



Mark Scheme (Results)

Summer 2017

Pearson Edexcel International GCSE
in Chemistry (4CH0) Paper 1CR

Pearson Edexcel International
in Science (Double Award) (4SC0) Paper 1CR

Edexcel and BTEC Qualifications

Edexcel and BTEC qualifications are awarded by Pearson, the UK's largest awarding body. We provide a wide range of qualifications including academic, vocational, occupational and specific programmes for employers. For further information visit our qualifications websites at www.edexcel.com or www.btec.co.uk. Alternatively, you can get in touch with us using the details on our contact us page at www.edexcel.com/contactus.

Pearson: helping people progress, everywhere

Pearson aspires to be the world's leading learning company. Our aim is to help everyone progress in their lives through education. We believe in every kind of learning, for all kinds of people, wherever they are in the world. We've been involved in education for over 150 years, and by working across 70 countries, in 100 languages, we have built an international reputation for our commitment to high standards and raising achievement through innovation in education. Find out more about how we can help you and your students at: www.pearson.com/uk

Summer 2017

Publications Code 4CH0_1CR_1706_MS

All the material in this publication is copyright

© Pearson Education Ltd 2017

General Marking Guidance

- All candidates must receive the same treatment. Examiners must mark the first candidate in exactly the same way as they mark the last.
- Mark schemes should be applied positively. Candidates must be rewarded for what they have shown they can do rather than penalised for omissions.
- Examiners should mark according to the mark scheme not according to their perception of where the grade boundaries may lie.
- There is no ceiling on achievement. All marks on the mark scheme should be used appropriately.
- All the marks on the mark scheme are designed to be awarded. Examiners should always award full marks if deserved, i.e. if the answer matches the mark scheme. Examiners should also be prepared to award zero marks if the candidate's response is not worthy of credit according to the mark scheme.
- Where some judgement is required, mark schemes will provide the principles by which marks will be awarded and exemplification may be limited.
- When examiners are in doubt regarding the application of the mark scheme to a candidate's response, the team leader must be consulted.
- Crossed out work should be marked UNLESS the candidate has replaced it with an alternative response.

Question number	Answer	Notes	Marks
1 (a)	<p>D (Br₂)</p> <p>The only correct answer is D</p> <p>A is not correct because Br is the symbol for bromine</p> <p>B is not correct because the 2 is a superscript not a subscript</p> <p>C is not correct because the 2 is not a subscript</p>		1
(b) (i)	<p>B (diffusion)</p> <p>The only correct answer is B</p> <p>A is not correct because condensation is the change of state from a gas to liquid</p> <p>C is not correct because evaporation is change of state from a liquid to gas</p> <p>D is not correct because the change of state from sublimation is solid to gas</p>		1
(ii)		ALLOW particles evaporate	2

	<p>M1 the bromine/liquid evaporates / the particles escape (from the liquid)</p> <p>M2 (bromine fills the gas jar because) the (gas) particles move freely/randomly/constantly</p>	<p>ALLOW (gas) particles spread</p> <p>ALLOW particles move from a high concentration to low (concentration)</p> <p>IGNORE references to diffusion</p> <p>ACCEPT molecules for particles</p> <p>REJECT atoms once only</p>	
(c)	<p>C</p> <p>The only correct answer is C</p> <p>A is not correct because NH_3 gas diffuses faster HCl gas</p> <p>B is not correct because NH_3 gas diffuses faster HCl gas</p> <p>D is not correct because the position indicated is too close to the right hand end of the tube</p>		1
		Total	5

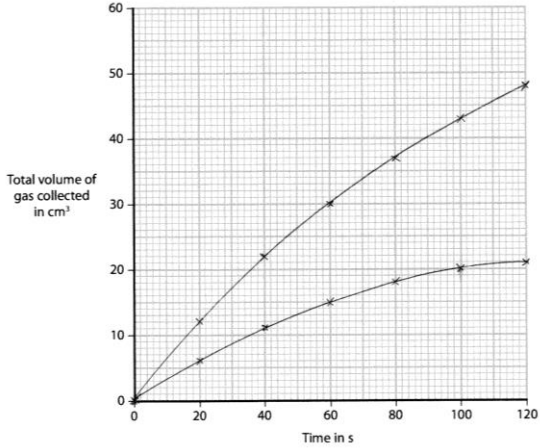
Question number	Answer	Notes	Marks								
2 (a)	<p>M1 oxygen</p> <p>M2 water (vapour)</p>	<p>ACCEPT O₂ IGNORE O IGNORE air</p> <p>ACCEPT moisture ACCEPT H₂O</p> <p>If both name and formula given, mark name only</p>	2								
(b)	<table border="1" data-bbox="575 651 1207 871"> <thead> <tr> <th data-bbox="575 651 862 687">Item</th> <th data-bbox="862 651 1207 687">Method</th> </tr> </thead> <tbody> <tr> <td data-bbox="575 687 862 724">bicycle chain</td> <td data-bbox="862 687 1207 724">oiling</td> </tr> <tr> <td data-bbox="575 724 862 799">bridge</td> <td data-bbox="862 724 1207 799">painting / galvanising</td> </tr> <tr> <td data-bbox="575 799 862 871">car body</td> <td data-bbox="862 799 1207 871">painting / galvanising</td> </tr> </tbody> </table>	Item	Method	bicycle chain	oiling	bridge	painting / galvanising	car body	painting / galvanising		3
Item	Method										
bicycle chain	oiling										
bridge	painting / galvanising										
car body	painting / galvanising										
(c)	<p>D (zinc)</p> <p>The only correct answer is D</p> <p>A is not correct because zinc is the only metal used to galvanise iron</p> <p>B is not correct because zinc is the only metal used to galvanise iron</p> <p>C is not correct because zinc is the only metal used to galvanise iron</p>		1								
		Total	6								

Question number	Answer	Notes	Marks										
3 (a)	<table border="1"> <thead> <tr> <th data-bbox="459 308 779 347">Separation</th> <th data-bbox="779 308 1088 347">Method</th> </tr> </thead> <tbody> <tr> <td data-bbox="459 347 779 456">to obtain sand from a mixture of sand and water</td> <td data-bbox="779 347 1088 456">filtration</td> </tr> <tr> <td data-bbox="459 456 779 564">to separate crude oil into its components</td> <td data-bbox="779 456 1088 564">fractional distillation</td> </tr> <tr> <td data-bbox="459 564 779 673">to obtain pure water from salt water</td> <td data-bbox="779 564 1088 673">simple distillation</td> </tr> <tr> <td data-bbox="459 673 779 782">to obtain ethanol from a mixture of ethanol and water</td> <td data-bbox="779 673 1088 782">fractional distillation</td> </tr> </tbody> </table>	Separation	Method	to obtain sand from a mixture of sand and water	filtration	to separate crude oil into its components	fractional distillation	to obtain pure water from salt water	simple distillation	to obtain ethanol from a mixture of ethanol and water	fractional distillation		4
Separation	Method												
to obtain sand from a mixture of sand and water	filtration												
to separate crude oil into its components	fractional distillation												
to obtain pure water from salt water	simple distillation												
to obtain ethanol from a mixture of ethanol and water	fractional distillation												
(b) (i)	<p>M1 (add to) <u>anhydrous/white</u> copper(II) sulfate</p> <p>M2 turns blue</p> <p>OR</p> <p>M1 add to cobalt(II) chloride paper / cobalt chloride paper</p> <p>M2 turns pink</p>	<p>ACCEPT turns copper(II) sulfate from white to blue for 2 marks</p> <p>ACCEPT blue cobalt(II) chloride ALLOW anhydrous cobalt(II) chloride</p> <p>ACCEPT turns cobalt(II) chloride from blue to pink for 2 marks</p> <p>M2 DEP on M1 or near miss e.g. just copper(II) sulfate IGNORE any reference to testing with indicators</p>	2										

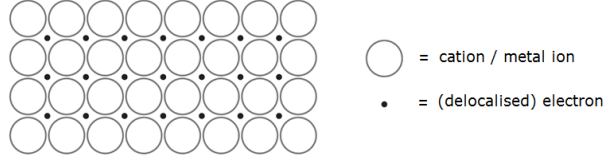
Question number	Answer	Notes	Marks
3 (b) (ii)	<p>M1 measure/determine/test its boiling point</p> <p>M2 100 °C</p> <p>OR</p> <p>M1 measure/determine/test its freezing point</p> <p>M2 0 °C</p> <p>OR</p> <p>M1 measure/determine/test its density</p> <p>M2 1 g/cm³</p>	<p>ACCEPT boil it / heat until it boils</p> <p>it boils at 100 °C</p> <p>ALLOW its boiling point is 100 °C for 1 mark</p> <p>ALLOW heat it and it boils at 100 °C for 2 marks</p> <p>ACCEPT freeze it / cool until it freezes</p> <p>it freezes at 0 °C</p> <p>ALLOW its freezing point is 0 °C for 1 mark</p> <p>ALLOW cool it and it freezes at 0 °C for 2 marks</p> <p>ALLOW its density is 1 g/cm³ for 1 mark</p> <p>M2 DEP on M1 throughout</p>	2
Total			8

Question number	Answer	Notes	Marks
4 (a) (i)	<p>B (the number of protons in an atom)</p> <p>The only correct answer is B</p> <p>A is not correct because atomic number is not the number of neutrons in an atom</p> <p>C is not correct because atomic number is not the number of protons plus the number of electrons in an atom</p> <p>D is not correct because atomic number is not the number of protons plus the number of neutrons in an atom</p>		1
	<p>(ii) C (electrons in the outer shell)</p> <p>The only correct answer is C</p> <p>A is not correct because the number of protons does not determine chemical properties</p> <p>B is not correct because the number of neutrons does not determine chemical properties</p> <p>D is not correct because the number of protons and neutrons does not determine chemical properties</p>		1

(b)		Li	Be	B	C	N	F	One mark for each correct row	3
	melting point				high	low	low		
	structure	giant			giant	molecular			
	acid-base character of the oxide	basic			acidic	acidic	acidic		
								Total	5

Question number	Answer	Notes	Marks
5 (a) (i)	add acid before magnesium	ORA	1
(ii)	a burette has a better resolution (than a measuring cylinder)	ALLOW greater accuracy (of data) ALLOW greater precision (of data)	1
(b)	 <p data-bbox="600 986 1265 1058">M1 and M2 all points plotted correctly to the nearest gridline for both experiments</p> <p data-bbox="600 1273 1279 1310">M3 suitable curve of best fit drawn for acid X</p> <p data-bbox="600 1345 1279 1382">M4 suitable curve of best fit drawn for acid Y</p>	<p data-bbox="1384 1066 1702 1241">Deduct one mark for each incorrectly plotted point Missing (0,0) loses 1 mark only</p>	4

Question number	Answer	Notes	Marks
5 (c)	<p>M1 Y (has the greater concentration)</p> <p>M2 (because) the curve (for acid Y) has a steeper slope/greater gradient (showing that the reaction is faster)</p> <p>OR (because) it produces the larger volume of gas/more gas in the same time</p>	<p>M1 DEP M2</p> <p>OWTTE</p> <p>Mark CSQ on candidate's labelling of their curves If no labelling assume steeper curve is acid Y</p>	2
(d)	<p>M1 vertical line drawn to touch curve at $t = 70$ s OR horizontal line drawn to touch curve at $t = 70$ s</p> <p>M2 value read correctly from candidate's graph to nearest gridline</p>	<p>Expected value in range 16-17 (cm^3)</p>	2
(e)	<p>M1 $17 \text{ (cm}^3\text{)}$</p> <p>M2 $(17 \div 30) = 0.57 \text{ (cm}^3\text{/s)}$</p>	<p>ACCEPT value read correctly from candidate's graph to nearest gridline</p> <p>ACCEPT any number of sig figs e.g. 0.6, 0.567, 0.56 recurring</p> <p>Mark M2 CSQ on M1</p>	2
		Total	12

Question number	Answer	Notes	Marks
6 (a) (i)	 <p> M1 minimum of 5 circles in regular pattern in at least two rows M2 circle(s) labelled cation(s)/metal ion(s)/positive ion(s)/atom(s) M3 spread of electrons in between circles AND labelled </p> <p> M1 delocalised electrons M2 are able to flow (in an electric field) </p>	<p> IGNORE different numbers of ions and electrons + drawn in circle is sufficient IGNORE particle e or e⁻ is sufficient </p> <p> ACCEPT sea of electrons IGNORE free electrons </p> <p> ACCEPT are able to move / are mobile IGNORE references to carrying a charge/current M2 DEP on mention of electrons Any mention of ions/atoms/nuclei/protons moving scores 0/2 </p>	<p>3</p> <p>2</p>

Question number	Answer	Notes	Marks
6 (a) (iii)	<p>M1 strong (electrostatic) forces (of attraction) between cations/metal ions and (delocalised) electrons</p> <p>M2 large amount of (thermal/heat) energy needed to overcome the forces</p>	<p>ACCEPT strong forces (of attraction) between nuclei of atoms and (delocalised) electrons</p> <p>ACCEPT strong metallic bonding / metallic bonds Not just strong bonds</p> <p>Not just heat Not just more energy</p> <p>M2 DEP on M1 or near miss, e.g. strong bonds</p> <p>Any mention of ionic bonds or covalent bonds or intermolecular forces being broken or overcome scores 0/2</p>	2

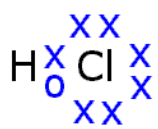
Question number	Answer	Notes	Marks
6 (b)	<p>M1 (Cu²⁺) blue precipitate</p> <p>M2 (Fe²⁺) green precipitate</p>	<p>IGNORE shades</p> <p>IGNORE names of precipitates or formulae</p> <p>REJECT any other colours</p> <p>Both colours correct but no mention of precipitates score 1/2</p>	2
(c) (i)	it is unreactive / it is not very reactive / it is low in the reactivity series	<p>ACCEPT it is not as reactive as iron/it is below iron in the reactivity series</p> <p>IGNORE inert</p>	1
(ii)	CO ₂ + C → 2CO / 2C + O ₂ → 2CO	ACCEPT multiples and halves	1
(iii)	<p>M1 iron / Fe</p> <p>M2 (because) it loses oxygen / because oxygen has been removed</p> <p>M2 DEP on M1</p>	<p>ALLOW it loses O</p> <p>IGNORE gains electrons</p> <p>ALLOW the iron(III) <u>ion</u> / Fe³⁺ (is reduced because it) has gained electrons for 1 mark</p> <p>ALLOW Iron(III) oxide/Fe₂O₃/Fe³⁺ (is reduced because it) has lost oxygen for 1 mark</p>	2
(iv)	<p>M1 calcium carbonate decomposes/forms/ changes into calcium oxide</p> <p>M2 calcium oxide reacts with silicon dioxide/silica</p>	<p>ACCEPT limestone for calcium carbonate</p> <p>ACCEPT correct chemical equations</p> <p>IGNORE sand</p> <p>ALLOW calcium carbonate reacts with silicon dioxide for 1 mark</p> <p>IGNORE reacts with impurities</p>	2
Total			15

Question number	Answer	Notes	Marks
7 (a) (i)	<p>silver does not react with (dilute sulfuric) acid</p>	<p>ACCEPT silver is below hydrogen in the reactivity series</p> <p>IGNORE silver is unreactive / silver has a low reactivity / silver is inert</p>	1
(ii)	<p>D (zinc and sulfuric acid)</p> <p>The only correct answer is D</p> <p>A is not correct because copper does not react with dilute sulfuric acid</p> <p>B is not correct because gold does not react with dilute hydrochloric acid</p> <p>C is not correct because the reaction between potassium and dilute hydrochloric acid is explosive and therefore not safe</p>		1
(b) (i)	<p>Experiment 3 because the volume collected is much lower than / very different to the other three</p>	<p>ACCEPT any answer that suggests the result/value in experiment 3 is much lower/much different to the other three e.g. it is much lower than the other three or there is a large difference between it and the other three</p>	1

<p>(ii)</p>	<p>M1 chooses 64, 67 and 63</p> <p>M2 65 (cm³)</p>	<p>ACCEPT 194</p> <p>ACCEPT any number of sig figs except 1 eg 64.7 / 64.67 / 64.667 / 64.6 recurring</p> <p>Correct answer with no working scores 2</p> <p>ALLOW 1 mark for correct calculation using all four volumes (58.5 / 59)</p>	<p>2</p>
<p>(iii)</p>	<p>to increase the validity/reliability (of the measurements/data)</p>	<p>ACCEPT to check for anomalous results</p> <p>IGNORE references to increased accuracy</p>	<p>1</p>
			<p>Total 6</p>

Question number	Answer	Notes	Marks
8 (a)	<p>M1 $(54.4 \div 127) = 0.428$</p> <p>AND</p> <p>$(45.6 \div 35.5) = 1.28$</p> <p>M2 Divide by the smaller number to obtain 1:3 ratio OR $0.428:1.28 = 1:3$</p>	<p>ACCEPT any number of sig figs except 1, but allow use of 0.4 in calculation of ratio in M2</p> <p>ALLOW answers to M1 given as fractions only if it is clear that division by smaller has taken place to obtain a ratio</p>	2
(b) (i)	<p>M1 rate of forward reaction = rate of backward reaction</p> <p>M2 amounts/concentrations/masses of reactants (and products) remain constant / constant macroscopic properties</p>	<p>IGNORE forward reaction = backward reaction</p> <p>IGNORE amounts/concentrations of reactants and products are equal</p>	2
(ii)	<p>M1 (liquid) (contains) ions that can flow/move/are mobile</p> <p>M2 (solid) (does not contain any) charged particles that can flow/move/are mobile</p>	<p>IGNORE references to carry charge REJECT any reference to electrons moving</p> <p>ACCEPT molecules are not charged/are neutral ACCEPT no electrons that can flow/move/are mobile ACCEPT no delocalised electrons ACCEPT no sea of electrons IGNORE free electrons REJECT any suggestion that the solid is ionic or contains ions</p>	2
		Total	6

Question number	Answer	Notes	Marks
9 (a) (i)	halogens are poisonous/toxic	ACCEPT any named halogen IGNORE harmful/dangerous/irritant IGNORE (named) products are toxic	1
(ii)	<p>M1 chlorine most reactive AND iodine least reactive</p> <p>OR</p> <p>chlorine > bromine > iodine</p> <p>M2 chlorine glows most brightly / glows very brightly</p> <p>AND</p> <p>iodine glows least brightly / glows dimly</p>	<p>IGNORE reactivity of the halogens decreases down the group</p> <p>IGNORE references to heat given out</p>	2
(iii)	<p>M1 the statement/student is incorrect</p> <p>M2 because vapours/gases were used (so the physical states at room temperature are irrelevant)</p>	<p>ACCEPT the reactivity can be found</p> <p>M1 DEP on M2</p>	2

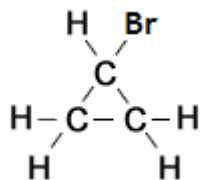
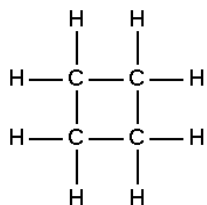
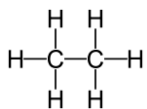
Question number	Answer	Notes	Marks
9 (b) (i)	$\text{H}_2 + \text{Br}_2 \rightarrow 2\text{HBr}$	ACCEPT multiples and halves	1
(ii)	 <p>M1 bonding pair of electrons</p> <p>M2 non-bonding pairs correct</p>	<p>M2 DEP on M1</p> <p>ALLOW any combination of dots and crosses</p> <p>If overlapping/touching circles used both electrons must be within the overlapping/touching area</p> <p>IGNORE inner electrons on chlorine even if incorrect</p>	2
(iii)	hydrochloric acid		1
Total			9

Question number	Answer	Notes	Marks
10 (a) (i)	<p>M1 (magnesium ribbon) shiny / silvery</p> <p>M2 (magnesium oxide) white (powder/solid/smoke/ash)</p>	<p>IGNORE grey</p> <p>IGNORE grey</p>	2
(ii)	<p>(lift the lid) to allow oxygen into the crucible</p> <p>(replaces the lid) minimise the loss of magnesium oxide/product</p>	<p>ACCEPT air</p> <p>ACCEPT to allow magnesium to react with <u>oxygen</u></p> <p>ACCEPT smoke for magnesium oxide</p> <p>ALLOW prevent the loss of magnesium oxide, etc</p>	2
(b)	<p>M1 $n[\text{Mg}] = 0.6 \div 24$ OR 0.025 (mol)</p> <p>M2 mass of $\text{O}_2 = 0.4$ (g)</p> <p>OR (M1 $\div 2$) $\times 32$ evaluated correctly</p> <p>Alternative method</p> <p>M1 48 (g) require 32 (g)</p> <p>M2 0.6 (g) require 0.4 (g)</p>		2
(c)	<p>$3\text{Mg} + \text{N}_2 \rightarrow \text{Mg}_3\text{N}_2$</p>	<p>ACCEPT multiples and halves</p>	1
Total			7

Question number	Answer	Notes	Marks
11 (a) (i)	<p>D (C_nH_{2n+2})</p> <p>The only correct answer is D</p> <p>A is not correct because C_nH_n is not the general formula for the alkanes</p> <p>B is not correct because C_nH_{2n-2} is not the general formula for the alkanes</p> <p>C is not correct because C_nH_{2n} is not the general formula for the alkanes</p>		1
	<p>(ii) C (C_nH_{2n})</p> <p>The only correct answer is C</p> <p>A is not correct because C_nH_n is not the general formula for the cycloalkanes</p> <p>B is not correct because C_nH_{2n-2} is not the general formula for the cycloalkanes</p> <p>D is not correct because C_nH_{2n} is not the general formula for the cycloalkanes</p>		1

(iii)	<p>Any two from:</p> <p>M1 similar/same chemical properties</p> <p>M2 graded physical properties / trend in physical properties</p> <p>M3 same functional group</p> <p>M4 each member differs (from the previous member) by CH₂</p>	<p>ALLOW they all react in the same way/in a similar way</p> <p>ACCEPT description of a graded physical property e.g. boiling increases as number of carbon atoms increases</p> <p>IGNORE different physical properties</p>	2
-------	---	--	---

Question number	Answer	Notes	Marks
11 (b) (i)	(the molecule) contains only single bonds	<p>ACCEPT contains no (carbon-carbon) double bonds/multiple bonds</p> <p>IGNORE references to no more atoms can be added, or contains the maximum number of hydrogen atoms</p>	1
	(ii)	IGNORE bond angles	1
	(iii)	IGNORE bond angles	1
(c) (i)	ultraviolet/uv (light/radiation)	<p>ALLOW sunlight</p> <p>IGNORE references to temperature or catalyst</p>	1
	(ii)	ACCEPT any number of bromine atoms substituted	1
Total			9



Question number	Answer	Notes	Marks
12 (a)	<p>M1 add water (and stir)</p> <p>M2 filter</p>	<p>ALLOW dissolve in water</p> <p>ACCEPT description of filtration</p> <p>ACCEPT decant</p> <p>M2 DEP on M1</p> <p>M2 not scored if any mention of evaporating the solution</p>	2
(b) (i)	<p>M1 (cation) ammonium / NH_4^+</p> <p>M2 (anion) chloride / Cl^-</p>	<p>If both name and formula given, both must be correct</p> <p>One mark if both correct but given in wrong order</p>	2
(ii)	ammonia / NH_3	If both name and formula given, both must be correct	1
(c) (i)	<p>M1 (anion) carbonate / CO_3^{2-}</p> <p>M2 (because) carbon dioxide/CO_2 is given off (when hydrochloric acid/HCl is added)</p>	<p>ACCEPT hydrogencarbonate / HCO_3^-</p> <p>If both name and formula given, both must be correct</p> <p>ALLOW the gas is carbon dioxide</p>	2
(ii)	<p>M1 (test) flame test</p> <p>M2 (result) brick-red (colour)</p>	<p>ACCEPT description of flame test</p> <p>ACCEPT red / orange-red</p> <p>REJECT all other colours</p> <p>M2 DEP on M1 or near miss e.g. heat the solid but REJECT if solid is heated in a test tube, etc</p>	2
Total			9

Question number	Answer	Notes	Marks
13 (a)	<p>M1 use the <u>burette</u> to add the (sulfuric) acid (to the sodium hydroxide)</p> <p>M2 until there is a change in colour (of the indicator/methyl orange/solution)</p> <p>M3 take initial and final readings of acid (and subtract to calculate the volume added)</p> <p>Plus any one from:</p> <p>M4 add acid dropwise (when near to the end point)</p> <p>M5 swirl the solution (when near to the end point)</p> <p>M6 repeat to obtain concordant results</p>	<p>If both initial and final colours are given both must be correct</p> <p>ACCEPT orange/pink/red as the final colour</p> <p>ACCEPT correct colours of any alternative indicator chosen e.g. (pink) to colourless for phenolphthalein</p> <p>(blue) to purple/red/pink for litmus</p> <p>REJECT Universal Indicator</p> <p>ALLOW repeat to obtain accurate/reliable results</p>	4

Question number	Answer	Notes	Marks
13 (b) (i)	20(.0) °C		1
	(ii) 17.5 cm ³		1
	(iii) 10 (cm ³) AND 25 (cm ³)		1
		Total	7

Question number	Answer	Notes	Marks
14 (a)	<p>B (changes from shiny to dull)</p> <p>The only correct answer is B</p> <p>A is not correct because a freshly exposed surface of lithium does not bubble and fizz when in contact with air</p> <p>C is not correct because a freshly exposed surface of lithium does not burst into flame when in contact with air</p> <p>D is not correct because a freshly exposed surface of does change when in contact with air</p>		1
(b) (i)	<p>burns with a pop/squeak</p> <p>OR</p> <p>use burning/lit spill to see if pops/squeaks</p> <p>OR</p> <p>use flame to see if pops/squeaks</p>	<p>Must be reference to test and result</p> <p>Reference to spill/match with no indication of flame is not enough</p> <p>REJECT reference to glowing spill/splint</p> <p>IGNORE flame extinguished</p> <p>'Squeaky pop test' alone is not sufficient</p>	1

(ii)	<p>M1 lithium hydroxide/LiOH/hydroxide <u>ion(s)</u>/OH⁻ (formed)</p> <p>M2 (therefore) the <u>solution</u> is alkaline</p>	<p>If both name and formula given both must be correct</p> <p>ACCEPT pH is of the <u>solution</u> greater than 7</p> <p>ALLOW <u>solution</u> is basic</p>	2
------	--	--	---

Question number	Answer				Notes	Marks												
14 (c)	<table border="1"> <thead> <tr> <th data-bbox="394 341 593 448">Name of compound</th> <th data-bbox="600 341 792 448">Formula of compound</th> <th data-bbox="799 341 1003 448">Formula of cation in compound</th> <th data-bbox="1010 341 1227 448">Formula of anion in compound</th> </tr> </thead> <tbody> <tr> <td data-bbox="394 453 593 517"></td> <td data-bbox="600 453 792 517">Li₂O</td> <td data-bbox="799 453 1003 517"></td> <td data-bbox="1010 453 1227 517"></td> </tr> <tr> <td data-bbox="394 521 593 585"></td> <td data-bbox="600 521 792 585"></td> <td data-bbox="799 521 1003 585"></td> <td data-bbox="1010 521 1227 585">CO₃²⁻</td> </tr> </tbody> </table>				Name of compound	Formula of compound	Formula of cation in compound	Formula of anion in compound		Li ₂ O						CO ₃ ²⁻		2
Name of compound	Formula of compound	Formula of cation in compound	Formula of anion in compound															
	Li ₂ O																	
			CO ₃ ²⁻															
14 (d) (i)	<table border="1"> <tbody> <tr> <td data-bbox="394 660 1144 716">the reaction with caesium is more vigorous</td> <td data-bbox="1151 660 1227 716">✓</td> </tr> <tr> <td data-bbox="394 721 1144 804">the reaction with caesium produces a different gas</td> <td data-bbox="1151 721 1227 804"></td> </tr> <tr> <td data-bbox="394 809 1144 892">the reaction with caesium produces an acidic solution</td> <td data-bbox="1151 809 1227 892"></td> </tr> <tr> <td data-bbox="394 896 1144 979">the reaction with caesium produces a different compound</td> <td data-bbox="1151 896 1227 979">✓</td> </tr> <tr> <td data-bbox="394 984 1144 1032">the reaction of caesium is endothermic</td> <td data-bbox="1151 984 1227 1032"></td> </tr> </tbody> </table>				the reaction with caesium is more vigorous	✓	the reaction with caesium produces a different gas		the reaction with caesium produces an acidic solution		the reaction with caesium produces a different compound	✓	the reaction of caesium is endothermic			2		
the reaction with caesium is more vigorous	✓																	
the reaction with caesium produces a different gas																		
the reaction with caesium produces an acidic solution																		
the reaction with caesium produces a different compound	✓																	
the reaction of caesium is endothermic																		
(d) (ii)	<p data-bbox="394 1169 837 1201">2Cs + 2H₂O → 2CsOH + H₂</p> <p data-bbox="394 1273 752 1305">M1 all formulae correct</p> <p data-bbox="394 1345 730 1377">M2 correctly balanced</p>				<p data-bbox="1261 1169 1700 1201">ACCEPT multiples and halves</p> <p data-bbox="1261 1345 1480 1377">M2 DEP on M1</p>	2												

Question number	Answer	Notes	Marks
15 (a)	<p>M1 $0.02(00) \times 0.2(00)$</p> <p>M2 $0.004(00)$ (mol)</p>	<p>ACCEPT 4 for 1 mark</p> <p>Correct answer with no working scores 2</p>	2
(b)	<p>M1 $0.004(00) \div 0.1(00)$</p> <p>OR</p> <p>M2 from (a) $\div 0.1(00)$</p> <p>M2 $0.04(00) \text{ dm}^3 / 40(.0) \text{ cm}^3$</p> <p>OR</p> <p>M2 from (a) $\div 0.1(00)$ correctly evaluated</p>	<p>Unit required</p> <p>Correct answer, using M2 from part (a), with no working scores 2</p>	2
(c)	<p>M1 $n(\text{NaOH}) = 0.03(00) \times 0.2(00)$ OR $0.006(00)$ (mol)</p> <p>M2 mass of NaOH = 0.24 g</p> <p>OR</p> <p>M1 $\times 40$ correctly evaluated</p>	<p>Correct answer with no working scores 2</p>	2
		Total	6

